Key:



1. A. What is the mass of **Carbon-12**? # protons \_\_\_\_\_\_\_ + # neutrons \_\_\_\_\_\_\_ = \_\_\_\_\_\_\_ amu or ­­­­­­­\_\_\_\_\_\_\_\_ g



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1. What is the mass of **Carbon-14**? # protons \_\_\_\_\_\_\_ + # neutrons \_\_\_\_\_\_\_ = \_\_\_\_\_\_\_ amu or ­­­­­­­\_\_\_\_\_\_\_\_ g



2. Why are electrons NOT represented in determine the mass of the atom in the above question?



Find the molar mass of the following **elements**:

For example, the molar mass of Magnesium is 24.31g.



1. Calcium \_\_\_\_\_\_\_\_\_\_\_ C. Lithium \_\_\_\_\_\_\_\_\_\_\_\_ E. Barium \_\_\_\_\_\_\_\_\_\_\_



1. Chlorine \_\_\_\_\_\_\_\_\_\_\_ D. Oxygen \_\_\_\_\_\_\_\_\_\_\_\_ F. Fluorine \_\_\_\_\_\_\_\_\_\_\_



Find the molar mass of the following **compounds**:

For example, the molar mass of Magnesium Oxide, or MgO, would be 24.31 + 16.00 = 40.31g

& Sodium Oxide Na2O is (2 X 22.99) + 16.00 = 61.98g



1. CaCl2 \_\_\_\_\_\_\_\_\_\_\_ C. Li2O \_\_\_\_\_\_\_\_\_\_\_\_ E. BaF2 \_\_\_\_\_\_\_\_\_\_\_\_



1. LiCl \_\_\_\_\_\_\_\_\_\_\_ D. CaO \_\_\_\_\_\_\_\_\_\_\_\_ F. Ba(ClO3)2 \_\_\_\_\_\_\_\_\_\_\_

